



Most reactive



Potassium

Sodium

Calcium

Magnesium

Aluminium

Carbon

Zinc

Iron

Tin

Lead

Copper

Silver

Gold

Platinum

Least reactive

1. Keyword

Reactivity	How easily a substance takes part in a chemical reaction
Acidic	pH value less than 7
Alkaline	pH value more than 7
Oxide	Compound containing oxygen and another element
Displacement reaction	Where a more reactive element takes the place of a less reactive element in a compound

2. Reactions of Metals and Acids

Metal + acid → a salt + hydrogen
 The reaction between metal and acid gets faster when more reactive metals are used.

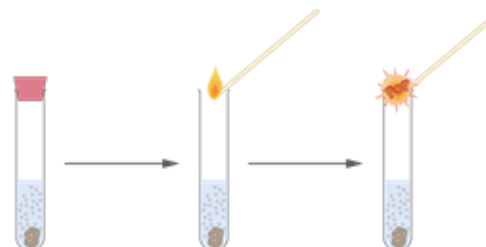
3. Reactions of Metals and water

Alkali Metal + Water → Alkali metal hydroxide + hydrogen

 Calcium, Magnesium, zinc, iron and lead react with steam to form a metal oxide and hydrogen gas
 E.g. Calcium + Water → Calcium oxide + hydrogen

4. Test for Hydrogen

Lit splint
 Makes a squeaky pop sound



5. Metal oxides

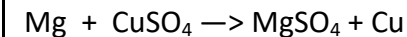
Bases – they dissolve to form alkaline solutions

Non-metal oxides

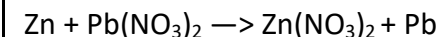
They dissolve in water to form acidic solutions

6. Displacement Reactions

Magnesium + copper sulphate → magnesium sulphate + copper



Zinc + lead nitrate → zinc nitrate + lead



The more reactive metal replaces the less reactive metal

7. Extraction of Metal

less reactive than carbon:
 Extracted from their metal oxide by carbon.
 Metal oxide + carbon → metal + carbon dioxide
 More reactive than carbon:
 Extracted from their metal oxide by electrolysis
 Gold, Silver and Platinum found in their native state (they are unreactive)